



COURSE OUTLINE: CHM190 - CHEMISTRY I

Prepared: Christine Giardino

Approved: Bob Chapman, Chair, Health

Course Code: Title	CHM190: CHEMISTRY I FOR PADD
Program Number: Name	3065: PRE-HEALTH DIP DGR
Department:	PRE-HEALTH
Semesters/Terms:	20F, 21W
Course Description:	In this course, students will examine the fundamental concepts, procedures, and calculations of chemistry. Course work will include examples and problems that relate to health and the human body. Topics in this course include properties of matter, chemical bonding, atomic and molecular structure, chemical nomenclature, chemical quantities, chemical equations, stoichiometry, the gas laws, and solutions and solubility. Laboratory investigations in this course will focus on safety, measurement, and common practices and procedures. The purpose of the lab work is to develop practical skills while gaining a better understanding of the theoretical concepts and calculations.
Total Credits:	5
Hours/Week:	5
Total Hours:	75
Prerequisites:	There are no pre-requisites for this course.
Corequisites:	There are no co-requisites for this course.
This course is a pre-requisite for:	CHM191
Vocational Learning Outcomes (VLO's) addressed in this course: Please refer to program web page for a complete listing of program outcomes where applicable.	3065 - PRE-HEALTH DIP DGR VLO 2 Examine concepts, processes and systems of chemistry, including atomic and molecular structure; quantities in chemical reactions; solutions and solubility; acids and bases; as well as organic chemistry and biochemistry in relation to health and the human body. VLO 6 Investigate health sciences and science-related questions, problems and evidence using the scientific method.
Essential Employability Skills (EES) addressed in this course:	EES 3 Execute mathematical operations accurately. EES 4 Apply a systematic approach to solve problems. EES 5 Use a variety of thinking skills to anticipate and solve problems. EES 6 Locate, select, organize, and document information using appropriate technology and information systems. EES 8 Show respect for the diverse opinions, values, belief systems, and contributions of others.

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	EES 9 Interact with others in groups or teams that contribute to effective working relationships and the achievement of goals. EES 10 Manage the use of time and other resources to complete projects. EES 11 Take responsibility for ones own actions, decisions, and consequences.								
General Education Themes:	Science and Technology								
Course Evaluation:	Passing Grade: 50%, D A minimum program GPA of 2.0 or higher where program specific standards exist is required for graduation.								
Other Course Evaluation & Assessment Requirements:	Students must achieve an average of 50% on tests, independent of other components, to obtain a passing grade in this course. Students must complete all labs designated as mandatory to obtain a passing grade in this course.								
Books and Required Resources:	General, Organic, and Biological Chemistry: Structures of Life by Karen C. Timberlake Publisher: Pearson Edition: 6 ISBN: 9780134730684 or 9780134763071 Beyond Labz purchase software download from www.beyondlabz.com								
Course Outcomes and Learning Objectives:	<table border="1"> <thead> <tr> <th>Course Outcome 1</th><th>Learning Objectives for Course Outcome 1</th></tr> </thead> <tbody> <tr> <td>1. Perform measurements and calculations that are accurate and precise.</td><td> 1.1 Identify a number as measured or exact. 1.2 Indicate the uncertainty associated with a particular measurement. 1.3 Use appropriate metric or SI units, especially when recording measurements of length, mass, volume, temperature, and time. 1.4 Express a measured or calculated value using scientific notation. 1.5 Round off the result of a calculation to the appropriate number of significant digits. </td></tr> <tr> <th>Course Outcome 2</th><th>Learning Objectives for Course Outcome 2</th></tr> <tr> <td>2. Classify matter and describe its structure, organization, and properties.</td><td> 2.1 Classify matter based on its state (solid, liquid, or gas) and composition (type of pure substance or mixture) using appropriate terminology. 2.2 Distinguish between physical and chemical properties. 2.3 Describe atomic structure and determine numbers of subatomic particles in a given isotope. 2.4 Define isotope and explain the effect that isotopes have on the atomic mass of an element. 2.5 Represent an isotope using its atomic and mass numbers. 2.6 Use the octet rule to determine the charge of an ion. 2.7 Apply modern atomic theory to determine the electron configuration for an element. </td></tr> </tbody> </table>	Course Outcome 1	Learning Objectives for Course Outcome 1	1. Perform measurements and calculations that are accurate and precise.	1.1 Identify a number as measured or exact. 1.2 Indicate the uncertainty associated with a particular measurement. 1.3 Use appropriate metric or SI units, especially when recording measurements of length, mass, volume, temperature, and time. 1.4 Express a measured or calculated value using scientific notation. 1.5 Round off the result of a calculation to the appropriate number of significant digits.	Course Outcome 2	Learning Objectives for Course Outcome 2	2. Classify matter and describe its structure, organization, and properties.	2.1 Classify matter based on its state (solid, liquid, or gas) and composition (type of pure substance or mixture) using appropriate terminology. 2.2 Distinguish between physical and chemical properties. 2.3 Describe atomic structure and determine numbers of subatomic particles in a given isotope. 2.4 Define isotope and explain the effect that isotopes have on the atomic mass of an element. 2.5 Represent an isotope using its atomic and mass numbers. 2.6 Use the octet rule to determine the charge of an ion. 2.7 Apply modern atomic theory to determine the electron configuration for an element.
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	2.8 Relate an element's electron configuration to its position (s, p, or d block) on the periodic table. 2.9 Recognize and compare the properties of elements in subdivisions of the periodic table: periods, groups, metals, non-metals, and metalloids. 2.10 Identify and explain periodic patterns and trends with respect to atomic size, ionization energy, electronegativity, number of energy levels, number of valence electrons, and reactivity.
Course Outcome 3	Learning Objectives for Course Outcome 3
3. Explain and apply principles of chemical bonding.	3.1 Determine the number of valence electrons and their involvement in the formation of chemical bonds. 3.2 Determine the polar character of bonds using the concept of electronegativity. 3.3 Write the Lewis structures for atoms, molecules and polyatomic ions. 3.4 Apply VSEPR theory to draw linear, trigonal planar, trigonal pyramidal, tetrahedral, and bent shapes for covalently bonded molecules. 3.5 Use electronegativity and molecular shape to determine the polarity of a molecule. 3.6 Relate polarity to the forces (i.e. hydrogen bonds, dipole-dipole attractions, and dispersion forces) that occur between molecules. 3.7 Explain how bond types and intermolecular forces account for the properties of molecular and ionic compounds, covalent networks, and metallic substances.
Course Outcome 4	Learning Objectives for Course Outcome 4
4. Represent ions, diatomic molecules, and compounds by name and chemical formula.	4.1 Use the octet rule to predict ion symbols and chemical formulae. 4.2 Use the chemical formula of a compound to determine its IUPAC name. 4.3 Use the IUPAC name of a compound to determine its chemical formula.
Course Outcome 5	Learning Objectives for Course Outcome 5
5. Perform calculations and conversions involving chemical quantities, temperature, and energy.	5.1 Define and determine the mass, volume, density, or specific gravity of a substance. 5.2 Define and explain the meaning of moles, mass, molar mass, and Avogadro's number. 5.3 Calculate moles, mass, molar mass, and number of particles for a substance. 5.4 Determine the percent composition of a compound or mixture. 5.5 Describe three commonly used temperature scales and perform conversions from one to another. 5.6 State commonly used energy units and perform conversions from one to another.
Course Outcome 6	Learning Objectives for Course Outcome 6

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	6. Write balanced chemical equations and analyze reactions both qualitatively and quantitatively.	6.1 Characterize and distinguish between physical and chemical changes. 6.2 Identify the parts of an equation representing a physical or chemical change. 6.3 Predict the products and write balanced equations for combination, decomposition, combustion, single replacement, and double replacement reactions. 6.4 Identify a reaction as exothermic or endothermic. 6.5 Identify and describe the stoichiometric relationships among all of the chemical species involved in a balanced chemical reaction. 6.6 Explain how the following are related: coefficients in a balanced chemical equation, quantity in moles, mass, and number of particles. 6.7 Given a quantity in moles or mass of a substance, use a mole-mole factor from the balanced equation to calculate the moles or mass of another substance in the reaction. 6.8 For a given reaction, determine the limiting reactant and calculate the percent yield of a product. 6.9 Write a net ionic equation for a reaction occurring in aqueous solution. 6.10 Identify a redox reaction and determine which reactant is oxidized and which reactant is reduced. 6.11 Label the parts of an electrochemical cell and use half-reactions to explain electron transfer.
	Course Outcome 7	Learning Objectives for Course Outcome 7
	7. Describe the types and properties of solutions, the concept of concentration, and perform calculations involving solutions.	7.1 Use appropriate terminology to describe the types and properties of solutions. 7.2 Explain the importance of water as a solvent. 7.3 Describe the factors that affect the solubility of a solute. 7.4 Explain and predict the solubility of both ionic and molecular compounds. 7.5 Use solubility data to calculate the quantity of solute that will dissolve in a given amount of solvent. 7.6 Define and differentiate between strong, weak, and non-electrolytes. 7.7 Define concentration and perform calculations involving mass/mass percent, mass/volume percent, volume/volume percent, and molarity for ions and compounds in solution. 7.8 Describe the dilution process and perform dilution calculations. 7.9 Perform stoichiometric calculations involving reactions in solution. 7.10 Describe and differentiate between diffusion and osmosis.
	Course Outcome 8	Learning Objectives for Course Outcome 8
	8. Describe the properties of gases and apply the gas laws qualitatively and quantitatively.	8.1 Use appropriate terminology to describe the properties of an ideal gas. 8.2 Explain and apply Avogadro's Law, Boyle's Law, Charles' Law, Gay-Lussac's Law, the Combined Gas Law, and the Ideal Gas Law to real-world and health-related situations.

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		8.3 Solve quantitative problems involving the gas laws. 8.4 Perform stoichiometric calculations for chemical reactions involving gases.						
	Course Outcome 9	Learning Objectives for Course Outcome 9						
	9. Conduct laboratory investigations using appropriate scientific techniques.	9.1 Follow a provided lab procedure safely, accurately and efficiently. 9.2 Select the most appropriate equipment for performing a particular measurement. 9.3 Use proper measurement techniques for the precise and accurate collection of quantitative data. 9.4 Make clear and concise qualitative observations. 9.5 Record data using tables. 9.6 Use scientific reasoning to draw conclusions that explain investigation results.						
Evaluation Process and Grading System:	<table><tr><th>Evaluation Type</th><th>Evaluation Weight</th></tr><tr><td>Lab Assignments</td><td>25%</td></tr><tr><td>Tests</td><td>75%</td></tr></table>		Evaluation Type	Evaluation Weight	Lab Assignments	25%	Tests	75%
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Date:	August 26, 2020							
Addendum:	Please refer to the course outline addendum on the Learning Management System for further information.							

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